

Write out the steps for balancing chemical reactions. Include some of the "shortcuts".

Complete p. 239 #31 in your textbook.

Complete p. 337 #21

List seven types of chemical reactions (include dissolving). Provide a skeletal equation and an example for each one.

Construct a table of the three types of intermolecular forces (IMFs) and their characteristics; include changes in boiling points, density, and IMF strength. Give examples of molecules which exhibit each force.

List the steps for determining the % ionic character of a single bond

Complete p.349 #70. (If you have trouble, see your instructor)

Complete p.331 #13 and 14

Complete p.339 #24 (all)

Make a table showing the five major molecular shapes; include the bond angles, the number of bonding pairs on the central atom, and the number of non-bonding pairs on the central atom. Also give an example of each one.

Make a table showing the four types of molecular polarities. Give an example of each.

Explain how the polarity of a molecule is determined

Explain what a dipole is and provide an example

Draw a  $\Delta EN$  line (0.0 to 4.0) indicating the different types of bond polarity. Include an example of each one.

Complete p.247 #61.

Explain the concept of temporary dipoles in non-polar molecules

List the 6 rules of symmetry for determining the polarity of a molecule

Complete p.248 #64 b.

What is VSEPR theory and what do the letters stand for?

Complete p.244 #37 (b, c, and d) in your textbook.