

Describe at least 3 characteristics each of metals, metalloids, and non-metals.

Complete #23 on page 270 of your textbook.

Complete #61 on page 281 of your textbook

Discuss the difference between an anion and a cation; explain what happens to the radius of an atom when it loses or gains electrons.

Complete p.239 #30

Complete p.247 #46.

Explain why noble gases are stable.

Explain how Lewis dot structures are drawn for ionic bonds; provide an example.

Draw the Lewis dot structures for the ionic compounds: Na_3P and MgF_2 .

Explain what a metallic lattice is; make sure to use the term "sea of electrons"

Draw a periodic table with arrows indicating how the following trends change: atomic radius, ionization energy, electronegativity, and metallic properties.

Draw a periodic table. Label each of the families; include the transition and inner transition metals; indicate where the metals, metalloids, and non-metals are.

Complete p. 199 #18.

Make a table with the following: Family name, # of valence electrons, lose/gain electrons, charge on ion, and example of a dot structure. Make sure to include the transition metals.

Create a flowchart depicting how to name ionic and covalent compounds. Include the use of prefixes and Roman numerals where appropriate. List prefixes up to ten.

Explain why atomic radius decreases as one goes left to right across a period, but increases down a column.

Explain how Lewis dot structures are drawn for covalent bonds; provide an example with single bonds and another example with a double or triple bond

Complete p.199 #19.

Draw a Δ EN line indicating the different types of bond polarity. Include an example of each one.

Complete practice problems #10 and #11 on page 263 of your text.

Describe the octet rule and why it occurs.

Describe at least two differences between ionic and covalent compounds.

Complete p.207 #42.

Explain the relationship between electronegativity, ionization energy, and valence electrons

Complete p.279 #39.