

Differentiate between physical and chemical **changes**, and give as many examples as you can fit into this box.

Differentiate between physical and chemical **properties**, and give as many examples as you can fit into this box.

Explain the difference between ionic compounds and covalent compounds. Provide at least 2 examples of each.

Complete Practice Problem #24 on p.117 of your textbook. Show all work and calculations. Make sure to pay attention to significant figures.

Explain what an a.m.u. is, how it's defined, and how many grams it is.

Describe the goal(s) of the alchemists.

List all the elements which exist as molecular with their correct formulas; identify which ones are diatomic.

Provide a table with the following information about protons, electrons, and neutrons: charge, symbol, location, and mass (in grams).

List 5 clues that a chemical change has occurred.

Complete Understanding Concepts #65 on p.123 of your text. Show all work and calculations. Make sure to pay attention to significant figures.

Using a beautiful diagram and a detailed explanation, describe Rutherford's Gold Foil experiment and the conclusions he was able to make as a result.

Describe 2 characteristics, 2 types, and 2 examples of each of the following: elements, compounds, and mixtures.

Use a diagram to model six atomic models: Greek Model, Dalton's Model, Thomson's Model, Rutherford's Model, the Bohr Model, and the Modern Wave Model. Provide a quick labeled sketch. Also provide a brief description which distinguishes each model from the others.

List the difference between intrinsic properties and extrinsic properties. Provide lots of examples of each.

Write the complete chemical symbol for the ion with 36 electrons, 34 protons, and 45 neutrons.

Complete Practice Problem #23 on p.117 of your textbook. Show all work and calculations. Make sure to pay attention to significant figures.

Explain the difference between the atomic number, the mass number, and the atomic weight/mass

Complete all of #81 on page 60 of your textbook.

Complete #1 thru #6 on page 125, and #1 thru #5 on page 61. Place your answers here.

Contrast homogeneous and heterogeneous mixtures. Give an example of each.

Write the complete chemical symbol for the ion with 31 neutrons, 26 electrons, and 28 protons.

Make a diagram of a neutral carbon atom with the correct number of protons, neutrons, and electrons. Label the subatomic particles correctly.

Write the complete chemical symbol for the ion with 31 protons, 39 neutrons, and 28 electrons.

Explain the difference between a cation and an anion. Give an example of each.

Give the symbols, names, and numbers of subatomic particles for the three isotopes of hydrogen.