

Write out the steps for balancing chemical reactions. Include some of the "shortcuts".

Complete p.298 #17. Show all work and units.

Determine the molar masses of:

H₂O

CO₂

O₂

H₂

methane

ammonia

Complete p.291 #4. Show all work and units.

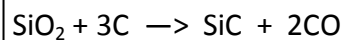
Complete p.299 #19. Show all work and units.

Complete p.292 #6. Show all work and units.

Complete p.371 #28. Show all work and units. Also identify the limiting reactant.

List seven types of chemical reactions (include dissolving). Provide a skeletal equation and an example for each one.

Complete p.349 #70. (If you have trouble, see your instructor)



If 50.0 g SiO₂ are used, what is the theoretical yield of SiC?

If only 32.2 g of SiC are actually produced, what is the percent yield?

Explain what an a.m.u. is and how it is derived. Include its mass.

Create a flowchart depicting how to name ionic and covalent compounds. Include the use of prefixes and Roman numerals where appropriate. List prefixes up to ten.

Draw the map of Mole Country.
Include **EVERYTHING**.

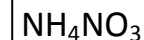
Explain the difference between a limiting reagent and an excess reagent.

Explain the difference between an empirical formula and a molecular formula; provide examples of each.

Complete p.360 #12 (a and b). Show all work and units. (mole-mole)

Complete p.364 #15. Show all work and units. (mass-molecules)

Calculate the percent composition of nitrogen in the fertilizers:



Complete p.379 #41 (b). Show all work and units. (mole-mass)

Complete p.361 #13. Show all work and units. (mass-mass)

Complete p.365 #17 and #18. Show all work and units. (volume-volume)

List 5 quantities a balanced chemical equation can be interpreted in terms of

Write the formulas for percent yield, percent error, and percent composition.

Put Avogadro's number here. In large print.