

COMMON ION TABLE

CATIONS (+)

+1 charge

Na ⁺¹	sodium
K ⁺¹	potassium
H ⁺¹	hydrogen
Li ⁺¹	lithium
Cu ⁺¹	copper (I)
Ag ⁺¹	silver
Au ⁺¹	gold (I)

+2 charge

Mg ⁺²	magnesium
Ca ⁺²	calcium
Ba ⁺²	barium
Zn ⁺²	zinc
Hg ⁺²	mercury (II)
Cu ⁺²	copper (II)
Pb ⁺²	lead (II)
Fe ⁺²	iron (II)
Cr ⁺²	chromium (II)
Mn ⁺²	manganese (II)

+3 charge

Al ⁺³	aluminum
Fe ⁺³	iron (III)
Cr ⁺³	chromium (III)
Au ⁺³	gold (III)

+1 charge (polyatomic)

NH ₄ ⁺¹	ammonium
H ₃ O ⁺¹	hydronium

+4 charge

Pb ⁺⁴	lead (IV)
Mn ⁺⁴	manganese (IV)

ANIONS (-)

-1 charge

F ⁻¹	fluoride
Cl ⁻¹	chloride
Br ⁻¹	bromide
I ⁻¹	iodide
H ⁻¹	hydride (rare)

-2 charge

O ⁻²	oxide
S ⁻²	sulfide

-3 charge

N ⁻³	nitride
P ⁻³	phosphide

-1 charge (polyatomic)

NO ₃ ⁻¹	nitrate
NO ₂ ⁻¹	nitrite
OH ⁻¹	hydroxide
CH ₃ COO ⁻¹	acetate
C ₂ H ₃ O ₂ ⁻¹	also acetate!
HCO ₃ ⁻¹	bicarbonate
ClO ₃ ⁻¹	chlorate
ClO ₄ ⁻¹	perchlorate
ClO ₂ ⁻¹	chlorite
ClO ⁻¹	hypochlorite
CN ⁻¹	cyanide
SCN ⁻¹	thiocyanate
MnO ₄ ⁻¹	permanganate

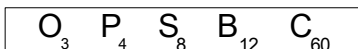
-2 charge (polyatomic)

SO ₄ ⁻²	sulfate
SO ₃ ⁻²	sulfite
CrO ₄ ⁻²	chromate
Cr ₂ O ₇ ⁻²	dichromate
CO ₃ ⁻²	carbonate
C ₂ O ₄ ⁻²	oxalate
O ₂ ⁻²	peroxide
S ₂ O ₃ ⁻²	thiosulfate

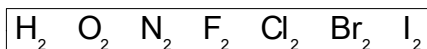
-3 charge (polyatomic)

PO ₄ ⁻³	phosphate
PO ₃ ⁻³	phosphite
BO ₃ ⁻³	borate

MOLECULAR ELEMENTS



DIATOMIC MOLECULAR ELEMENTS



COMMON MOLECULES TO KNOW

H ₂ O	water
NH ₃	ammonia
CH ₄	methane
CO ₂	carbon dioxide
SiO ₂	sand, glass
C ₆ H ₆	benzene
NaCl	table salt